



## Exercise 5

Date:

1. What does  $\delta^+$  mean? Where is it used?

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2. Complete the following table.

Type of bond	Constituent particles	How does the bond take place?
2.1	2.2	Positive atomic cores are bonded to a communal sea of valence electrons.
2.3	Molecules	2.4
Ionic bond	2.5	2.6

3. What is a small covalent molecule?

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4. Which intermolecular force is found in the following circumstances?

	Description	Intermolecular force
4.1	The weakest intermolecular force	
4.2	Forces when potassium chloride dissolves in water.	
4.3	Forces between vinegar (acetic acid – $C_2H_5COOH$ ) and water	
4.4	Forces that allow oxygen to liquefy.	
4.5	Forces between ammonia molecules	



	Description	Intermolecular force
4.6	Between an ion and a polar molecule	
4.7	Between polar molecules and non-polar molecules	
4.8	Between two polar molecules	

5. Explain the difference between interatomic forces and intermolecular forces by completing the table below.

	Interatomic forces	Intermolecular forces
Type of force	Covalent forces	Dipole-dipole forces
Strength of force		
Sketch illustrating the force		